

Name	Small Test Tube	Large Test Tube
Mass CO ₂	162.13 - 161.75 = <u>.38 g CO₂</u>	205.70 - 205.08 = <u>.62 g CO₂</u>
Circumference (cm) C =	26.1 cm	28 cm
Radius (cm) r =		
Balloon (cm ³ or mL) V =		
Temp(K) T =	23°C + 273 = <u>296 K</u>	23°C + 273 = <u>296 K</u>
Pressure (mmHg) P =	766 mm Hg	766 mm Hg
Pressure (atm) P =		
Amount (mol) n =		
Molar Mass (g/mol)		

$$r = \frac{C}{2\pi}$$

$$V = \frac{4}{3}\pi r^3$$

$$n = \frac{PV}{RT}$$